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Activity Coefficients from Vapor-Liquid Interfaces: A Molecular Dynamics Approach for Separation Chemistry

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Context

- Understanding solvent extraction – Recalling of rare earth elements and spent nuclear fuels
- The equilibrium constant \( K^o \) describes the extraction process and is given by the Mass Action Law
- Activity coefficients
- A multiscale approach provides mesoscopic thermodynamic properties

Methods and concepts

- Measuring activity and the vapor pressure
- Solvent flow: Vapor pressure of the pure solvent always higher than that of a mixture
- Ideal case: Solvent vapor pressure by Raoult's law
- Otherwise: Aggregation increases, dissolution decreases solvent vapor pressure and thus activity

Aqueous dysprosium nitrate \( \text{Dy(NO}_3\text{)}_3 \) solutions

- Water activity \( a_W \)
- Osmotic coefficient \( \phi_W \)
- Activity coefficient \( Y_C \)

Organic phases – Binary Mixtures of DMDOHEMA and \( n \)-heptane

- Molecular dynamics simulation of liquid-vapor equilibria of organic solvent phases containing:
- Pure solvent
- Extractant
- Water
- Ions

Increasing computational cost

- Solvent activities in binary mixtures of \( n \)-heptane and DMDOHEMA show ideal behavior and follow Raoult's law in good agreement up to a high extractant concentration
- Attractive and repulsive interactions in the mixtures cancel out

Outlook and conclusion

- Aqueous salt solutions: Results from molecular dynamics simulation are in good agreement with experimental data for different nitrate salts. This approach can be used for the validation of force fields for MD simulation with respect to thermodynamic properties.
- Organic Phases: Method has been successfully applied on binary mixtures (\( n \)-heptane and DMDOHEMA, but also ethanol and water)\(^{(b)} \). Allows accessing more complex organic solvent phases containing water and ionic species.

- Simulation of activity coefficients of complex solutions with volatile solvents for understanding solvent extraction

References


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