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Cetyltrimethylammonium Bromide Silver Bromide Complex as the Capping Agent of Gold Nanorods

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A complex between cetyltrimethylammonium bromide (CTAB) surfactant and silver bromide (CTASB) is recognized by NMR and X-ray photoelectron spectroscopy (XPS) to be the entity at the surface of gold nanorods, resulting from an *in situ* formation in the classical scheme of synthesis. It can thus be introduced directly along with the initial reactants in place of silver(I) salt to produce a new effective synthesis of these objects. Complementary XPS and quartz crystal microbalance (QCM) measurements on macroscopic gold surfaces confirm a strong adsorption of CTASB that is higher than that of CTAB and any other CTAX surfactants. The role of CTASB as a rod inducing agent by surface complexation is stressed.

Introduction

One dimensional metallic nanostructures have aroused great attention due to their future potential applications, expected mainly from their size-dependent optical properties^{1–3} or their use as building bricks in bottom-up approaches.⁴ Biological applications such as detection or gene delivery are other targets of interest.⁵ Electrochemical deposition in templates,⁶ electrochemical synthesis in solution,⁷ photochemical synthesis,⁸ and the seed-mediated growth approach⁹ are nonexhaustive examples of the numerous existing methods for the synthesis of gold nanorods. Wet chemical methods are the preferred routes due to the cost-effective, large-scale production of nanorods, and among them the seed-mediated growth approach is particularly interesting due to the high nanorod yields that can be reached. Indeed, since the original work on seed-mediated synthesis of gold nanorods,⁹ extensive studies have been devoted to the increase of the yield, monodispersity, and stability of the particles. In fact, achieving control over the growth toward gold nanorods with high yields is a challenging task. High yields were reached through the use of silver nitrate salt along with the other additives.¹⁰ Shorter aspect ratio rods are obtained compared to those obtained without silver(I) nitrate. However, short aspect ratio rods are interesting in that they present both transverse and longitudinal intense plasmon bands in the visible region of the spectrum, making them promising candidates for sensing or imaging applications.^{11,12}

Several sources have been advanced as the driving force for anisotropy. It has been suggested¹³ that the presence of surfactant

bilayers around the nanorods promotes the development of anisotropy through a “zipping” process. Long chain surfactants, such as cetyltrimethylammonium bromide (CTAB), are indeed needed to ensure a high yield of nanorods with good anisotropic ratio, while short chain surfactants lead to a higher percent of spherical byproducts and/or shorter rods.¹⁴ The role of the surfactant counterion was also considered,^{15,16} concluding that only brominated long chain surfactants lead to significant yields of rod shaped particles. The final mandatory component to obtain high yield syntheses of rods is silver(I) nitrate.¹⁰ The most commonly cited role of silver(I) is a reaction with CTAB to form the silver bromide compound that would deposit onto some preferential gold rods’ facets and thus hinder their growth in that direction.¹⁵ However, it has also been suggested that metallic silver deposits preferentially onto certain faces of the rod, due to the more positive underpotential of deposition of silver(I) on (110) facets, thus poisoning these facets.¹⁷

In this study, we propose an alternative mechanism based on spectroscopic characterizations of a seedless variation of the classical route of synthesis using silver nitrate. We recognize the primordial role of an *in situ* formed complex of CTAB and silver bromide hereafter named CTASB which can also be prepared *ex-situ*.¹⁸ Accordingly, an improved synthesis involving the use of CTASB as a cosurfactant along with CTAB and without silver nitrate is performed, which leads to a high yield nanorod synthesis for which the anisotropy promoter is specifically identified.

Results and Discussion

In a first stage, gold nanorods were prepared via a seedless synthesis inspired from Jana:¹⁹ a 10 mL solution containing 3.5 mM aqueous tetrachloroauric acid, 0.2 M CTAB, and 1 mM silver nitrate was constituted. Next, 0.12 mL of 0.4 M ascorbic acid was added under stirring. Finally, 64 μ L of an ice-cold aqueous 1.6 mM sodium borohydride solution was added. The

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Table 1. ^1H NMR Chemical Shifts in ppm^a

protons	capped gold nanorods	CTASB	capped gold nanospheres	CTAB
$\alpha\text{-CH}_2$	3.31	3.32	3.38	3.44
$\text{N}(\text{CH}_3)_3$	3.12	3.12	3.15	3.19
$\beta\text{-CH}_2$	1.79	1.79	1.78	1.77
CH_2 main chain	1.30	1.30	1.29	1.30
CH_3 main chain	0.87	0.87	0.87	0.87
CH_2 main chain	does not appear	1.37	1.37	1.37

^a ^1H NMR spectra in the Supporting Information. Values reported for capped nanorods stand the same for the two ways of synthesis (initial presence of silver nitrate or CTASB).

reaction, followed using UV–visible spectroscopy, was complete in less than 20 min. Transmission electron microscopy (TEM) analysis of the final particles provides the extensive dimensions for the rods, that is, ~ 9 nm in diameter for 32 nm in length, with 90% rods along with very few spherical byproducts. As a reference, gold nanospheres were also prepared by the same synthesis method as above but simply without addition of silver nitrate.

Detailed information about the interaction between the surfactant and gold nanorods can be obtained from NMR measurements.²⁰ Table 1 presents the chemical shifts of the ^1H spectra for a D_2O solution of CTAB, as well as D_2O solutions of capped gold nanospheres and capped gold nanorods, extracted from the initial surfactant solution and washed in heavy water. For all the protons of the main chain, except the α ones, the chemical shifts are identical for all three solutions. On the contrary, the chemical shifts of $\alpha\text{-CH}_2$ and $\text{N}(\text{CH}_3)_3$ for the capped gold nanorods, gold nanospheres, and CTAB are different. Indeed, chemical shifts go slightly highfield from CTAB to capped gold nanospheres, probably due to the vicinity of the gold surface to the head's protons (Knight shift²¹). As for capped gold nanorods, the chemical shifts are even lower, indicative of an additional or a different contribution to the chemical shifts other than that of the Knight shift.

A CTAB silver complex (CTASB) was recently identified¹⁸ in the mixture of CTAB and silver nitrate. Accordingly, the ^1H NMR chemical shifts of CTASB are presented in Table 1. The chemical shifts for a solution of CTASB are indeed identical to the ones obtained for capped gold nanorods. The highfield chemical shifts for $\alpha\text{-CH}_2$ and $\text{N}(\text{CH}_3)_3$ protons as compared to solutions of CTAB and CTAB capped gold nanospheres can now be attributed to the presence of a silver atom near to the N-polar head protons. This influence is complementary to the simple Knight shift of gold observed on capped gold nanospheres. An enlargement of the peaks for capped nanorods protons from the headgroup is also observed, further supporting the surfactants' binding to the gold surface by their headgroup.

The preceding results suggest that CTASB, a stable complex formed *in situ*, could be the real promoter of the anisotropy of the nanoparticles. In order to confirm this hypothesis, a new synthesis scheme was designed. CTASB was first prepared *ex situ* as in ref 18. Its purity was checked by X-ray diffraction (XRD) and elemental analysis, providing its formula $\text{C}_{19}\text{H}_{42}\text{NAgBr}_2$. A 10 mL solution containing 3.5 mM aqueous tetrachloroauric acid, 0.2 M CTAB, and various concentrations (from 0.2 to 2 mM) of CTASB was then constituted. Next, to each solution, 0.12 mL of 0.4 M ascorbic acid was added under stirring. Finally, 64 μL of an ice-cold aqueous 1.6 mM

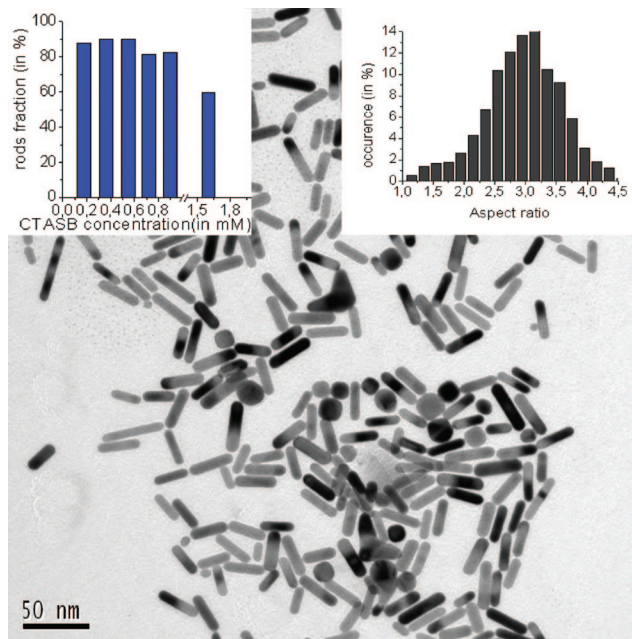


Figure 1. TEM picture of rods obtained with $[\text{CTASB}] = 0.55$ mM. Right inset: aspect ratio distribution from TEM images analysis. Left inset: rod fraction (percentage in number of objects) as a function of the CTASB initial concentration from TEM image analysis.

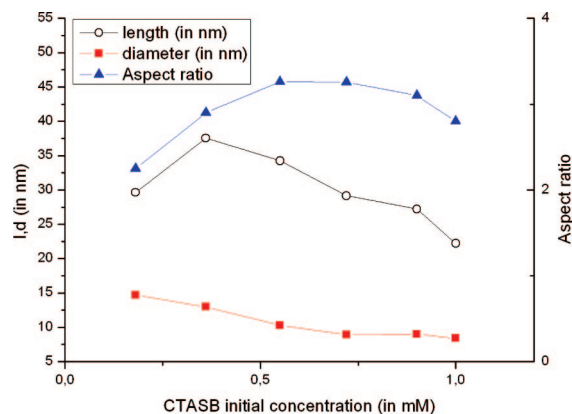


Figure 2. Dimensions of the gold nanorods as a function of the CTASB initial concentration.

sodium borohydride solution was added. This new scheme does not make use of silver(I) nitrate or another Ag^+ salt.

Using UV–visible measurements, the reaction was again followed for each CTASB concentration. The reaction is still complete in less than 20 min and provides nanorods. The XRD pattern of the gold nanorods shows that the particles are crystalline. Peaks are consistent with the bulk face-centered cubic (fcc) pure gold crystal. The respective intensities for the [111], [200], [220], and [311] peaks differ from the theoretical ones for bulk gold, indicating the preferential orientation along [111] and [100] planes in the gold nanorods and more to the point, among the two of them, the predominance of the [111] faces, that is, the side ones of the rods. The aspect ratio of the final particles, extracted from TEM analysis, reaches a maximum of 3.2 for 0.55 mM CTASB (Figure 1). A decrease in CTASB concentration leads to less anisotropic rods and an increase, to similar aspect ratios (ranging from 3.2 to 2.8) (Figure 2), but with the presence of more spheres (40% for 1.6 mM CTASB), as seen in the left inset in Figure 1. The evolution in the aspect ratios is confirmed by UV–visible measurements, though, as stated in ref 22, the estimated value

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Table 2. XPS Binding Energies in eV^a

element	capped gold nanorods	CTASB	capped gold nanospheres	CTAB	AgBr sol
Au 4f _{7/2}	84.0		84.0		
Au 4f _{5/2}	87.7		87.7		
Ag 3d _{5/2}	367.7	367.7 ¹⁸			367.7 ²⁷
Ag 3d _{3/2}	373.7	373.8 ¹⁸			373.7 ²⁷
Br 3d _{5/2}	67.8	67.9 ¹⁸	67.7	67.3 ²⁵ 67.5 ²⁶	68.3 ²⁷
N 1s	402.4(5)	402.3 ¹⁸ 402.4 ^b	402.1(5)	402 ²⁴ 402.1 ^b	

^a Nanorod Br 3d, N 1s, and Ag 3d XPS spectra in the Supporting Information. Values for nanorod synthesis in the presence of silver nitrate and CTASB stand the same. ^b This work.

is slightly different from the one obtained by TEM. According to TEM analysis, the highest rod length is reached for 0.4 mM CTASB, while for higher concentrations it decreases, from 37.5 to 22.5 nm (Figure 2). On the other hand, the diameter of the rods constantly decreases with an increase in CTASB, from 15 to 8 nm, suggesting the idea of a surface poisoning by CTASB. In conclusion, the optimum CTASB concentration is 0.55 mM, with more than 90% rods along with few spherical and cubical byproducts.

It was checked that the NMR spectrum of the extracted particles is the same as the one for the CTAB/silver nitrate synthesis.

In order to further characterize the CTASB species on the surface of the nanorods, X-ray photoelectron spectroscopy (XPS) measurements were conducted, on both CTAB/CTASB and CTAB/AgNO₃ nanorod syntheses as well as on pure CTASB. Several drops of the extracted nanoparticle solutions or a CTASB solution were deposited on a silicon wafer, dried under nitrogen, and then inserted in the XPS chamber. The XPS spectrum for gold nanorods showed the Au 4f_{7/2} and 4f_{5/2} doublet typical for Au(0) (Table 2). The Ag 3d_{3/2} and Ag 3d_{5/2} peaks are indicative of a Ag(I) oxidation state and stand the same as those for CTASB. The Br 3d_{5/2} binding energy for both gold nanorods samples is consistent with Br 3d_{5/2} in the pure CTASB spectrum and not with that of a silver bromide sol. The N1s peak can be attributed to N⁺.²³ Its binding energy at 402.45 eV for both gold nanorod samples is slightly higher than that for pure CTASB due to the gold surface vicinity. In fact, the vicinity of nitrogen with silver mainly explains the shift in energy for CTASB and capped gold nanorods as compared with CTAB whose N⁺ 1s peak is at ~402 eV.²⁴ Therefore, CTASB appears to be the only species present on the surface of the gold nanorods. The C 1s carbon peak for all samples was split into two contributions: one from the aliphatic carbons at 285 eV and another from the supposedly C–N ones at 286 eV. A Br/Ag ratio of 2 could be drawn from the XPS quantitative analysis of capped nanorods, with this value being in agreement with the chemical formula for CTASB. The XPS spectrum of gold nanospheres, synthesized without silver, was also taken. The binding energies for C 1s, Au 4f, and Br 3d present the same values as those for gold nanorods. However, the N⁺ 1s peak is now at 402.15 eV. This value is slightly higher than that for pure CTAB, indicative of the interaction between CTAB and the gold surface, but differs from the value for gold nanorods and CTASB. It thus confirms the capping of the gold nanorods by CTASB and not CTAB.

All these observations are indicative not only of an *in situ* formation of CTASB in the CTAB/silver nitrate synthesis but

also of a stronger adsorption energy for CTASB than that for CTAB, since after washing of the rods CTASB, though in lower concentration initially, as compared to CTAB, is the only species to be detected via XPS as well as ¹H NMR spectroscopy. In order to understand this exclusion of CTAB from the gold surface, the separate adsorption of these two surfactants onto a bulk gold surface was measured by quartz crystal microbalance (QCM) experiments. The gold covered quartz crystal, purchased from Maxtek, was exposed to a 1 mM CTASB solution and to a 1 mM CTAB solution. The adsorption of CTAB(SB) onto the gold is measured in terms of the frequency shift of the QCM (Figure S5, Supporting Information). The adsorbed amount was calculated with the Sauerbrey equation.²⁸ A total adsorbed mass of 150 ng of CTAB and 340 ng of CTASB was obtained, that is, 400 and 615 mmol, respectively, thus indicating that CTASB adsorbs around 1.5 more than CTAB onto a gold surface. As compared with previous QCM experiments led on several CTAX (X = NO₃⁻, OH⁻, Cl⁻, Br⁻, I⁻) surfactants¹⁶ where the maximum sorption on a gold surface was obtained with CTAB, CTASB appears to have an even increased adsorption to gold. It implies that when both CTAB (or another of the CTAX surfactants quoted above) and CTASB are present in the solution during the synthesis, the strong adsorption of CTASB implies a strong exclusion of CTAB or other CTAX surfactant from the gold surface and, all the more, that for each tested concentration the total amount of CTASB in solution is sufficient to cover the nanorod surface, even considering one CTASB molecule per gold surface atom. Further adsorption experiments were led on sputtered gold surfaces exposed to different surfactant solutions for 2 h and then thoroughly washed with water. The sputtered gold surfaces were then analyzed by XPS. Two groups could be observed. Gold surfaces exposed to a CTASB solution, CTAB/CTASB, or CTAB/AgNO₃ solutions with a CTAB concentration of 0.2 M and silver concentrations of 1.92 mM and 0.76 mM presented the C 1s, Ag 3d, N 1s, and Br 3d binding energies already reported for CTASB and CTASB capped gold nanorods. A gold surface exposed to a 0.2 M CTAB silverless solution has its N 1s peak at the same position as the one for CTAB gold nanospheres. For surfaces exposed to solutions of CTAB/CTASB or CTAB/AgNO₃ with a CTAB concentration of 0.2 M and a silver concentration of 0.48 mM, the N 1s peak appears to be at an intermediary value, at 402.3 eV, indicative of a mixed surface covering of both CTAB and CTASB. In that case, the concentration of CTASB is too low to dominate the surface adsorption. In addition, the quantification shows that the adsorption of bromide as well as nitrogen is increased in the presence of CTASB or silver nitrate as compared to the adsorption where the surfactant solution solely contains CTAB. Calculated atomic ratios between adsorption in the presence (0.76 or 1.92 mM) and absence of silver are ~1.8 and 2.9 for nitrogen and bromide, respectively. The final curvature of the rod interface must be due to a subtle balance between a strong adsorption of CTASB which wants to stabilize this Au–CTASB complex and the CTA chain association which promotes a directional growth effect.

The behavior of CTASB was studied by tensiometry, reporting it to have a surfactant behavior. However, no micelle formation could be detected, due to the CTASB low solubility limit in water (inferior to 1 mM). This limit explains why the synthesis cannot be led in a pure CTASB solution and has to make use

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of CTAB as a cosurfactant. In addition to the demonstrated CTASB surface complexation, a possible scenario is thus that CTAB acts as a chemical stabilizing agent for gold salt, through the formation of a $\text{CTA}^+ \text{AuBr}_4^-$ complex. In that case, it prevents a direct reduction in solution to metallic gold by ascorbic acid prior to NaBH_4 addition. Further studies along this direction are in progress.

Conclusion

We have set up a new way of synthesis of gold nanorods with the use of a complex cosurfactant along with CTAB. Silver is now inserted in the solution through this complex form. Spectroscopic characterizations of the final products have

evidenced the presence of the complex on the surface of the nanorods. Adsorption studies have confirmed the strong adsorption properties of CTASB on gold bulk surfaces and subsequently stressed its primordial role as a rod inducing agent.

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Supporting Information Available: NIR-visible, XPS, and ^1H NMR spectra of gold nanorods; XRD patterns of gold nanorods; and QCM adsorption measurements. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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